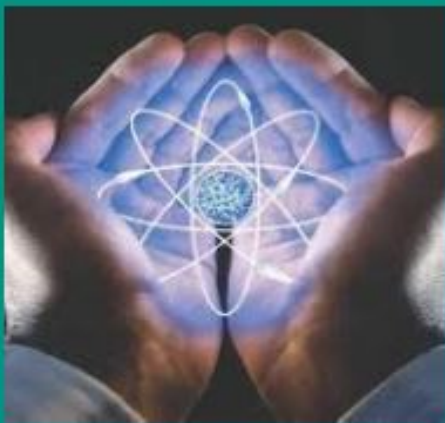


Table Of Content

Journal Cover	2
Author[s] Statement	3
Editorial Team	4
Article information	5
Check this article update (crossmark)	5
Check this article impact	5
Cite this article	5
Title page	6
Article Title	6
Author information	6
Abstract	6
Article content	7

Academia Open



By Universitas Muhammadiyah Sidoarjo

Originality Statement

The author[s] declare that this article is their own work and to the best of their knowledge it contains no materials previously published or written by another person, or substantial proportions of material which have been accepted for the published of any other published materials, except where due acknowledgement is made in the article. Any contribution made to the research by others, with whom author[s] have work, is explicitly acknowledged in the article.

Conflict of Interest Statement

The author[s] declare that this article was conducted in the absence of any commercial or financial relationships that could be construed as a potential conflict of interest.

Copyright Statement

Copyright © Author(s). This article is published under the Creative Commons Attribution (CC BY 4.0) licence. Anyone may reproduce, distribute, translate and create derivative works of this article (for both commercial and non-commercial purposes), subject to full attribution to the original publication and authors. The full terms of this licence may be seen at <http://creativecommons.org/licences/by/4.0/legalcode>

EDITORIAL TEAM

Editor in Chief

Mochammad Tanzil Multazam, Universitas Muhammadiyah Sidoarjo, Indonesia

Managing Editor

Bobur Sobirov, Samarkand Institute of Economics and Service, Uzbekistan

Editors

Fika Megawati, Universitas Muhammadiyah Sidoarjo, Indonesia

Mahardika Darmawan Kusuma Wardana, Universitas Muhammadiyah Sidoarjo, Indonesia

Wiwit Wahyu Wijayanti, Universitas Muhammadiyah Sidoarjo, Indonesia

Farkhod Abdurakhmonov, Silk Road International Tourism University, Uzbekistan

Dr. Hindarto, Universitas Muhammadiyah Sidoarjo, Indonesia

Evi Rinata, Universitas Muhammadiyah Sidoarjo, Indonesia

M Faisal Amir, Universitas Muhammadiyah Sidoarjo, Indonesia

Dr. Hana Catur Wahyuni, Universitas Muhammadiyah Sidoarjo, Indonesia

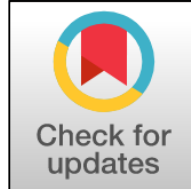
Complete list of editorial team ([link](#))

Complete list of indexing services for this journal ([link](#))

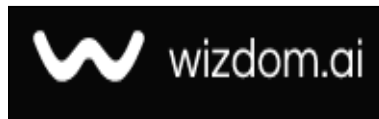
How to submit to this journal ([link](#))

Article information

Check this article update (crossmark)



Check this article impact ^(*)



Save this article to Mendeley



^(*) Time for indexing process is various, depends on indexing database platform

Method of Nickel Extraction from Industrial Waste and its Application in Production

Metode Ekstraksi Nikel dari Limbah Industri dan Penerapannya dalam Produksi

Bekhzod Farmanov Ilkhomovich, dildora@gmail.com, (1)

, Bangladesh

Tavashov Shakhzod Khujaxmatovich, dilbar.kuchkarova@mail.ru, (0)

, Uzbekistan

⁽¹⁾ Corresponding author

Abstract

The article describes the process of extracting nickel from the composition of the used waste of the GIAP-8 catalyst in the form of nickel nitrate, in which it not only does not affect the environmental environment, but is also a cost-effective solution for the production of nickel catalysts. The obtained catalysts were tested in laboratory conditions, the activity of the catalyst in this converted residual methane was 35,7% at T =500 °C and 6,5% at T= 700 °C.

Published date: 2022-06-30 00:00:00

Introduction

The development of modern science and technology leads to the fact that the need for non-ferrous and rare metals increases from year to year. Processing of non-ferrous metals leads to a reduction in deposits of rare earth metals. One of the urgent problems at present is the development and introduction into practice of energy-resource-saving technologies for processing rare metals contained in secondary raw materials into finished or semi-finished products. Rational integrated use of raw materials, transition to production are one of the most urgent problems of waste-free technologies of its processing and recycling of technogenic formations. Processing of raw materials and materials is of great economic and environmental importance.

According to numerous studies, the best catalyst for the conversion of methane is a nickel catalyst. And in the catalyst in the form of nickel oxide. The catalyst accelerates the conversion of methane. Therefore, in the process before starting the conversion, hydrogen together with the catalyst is reduced to nickel oxide, $\text{NiO} + \text{H}_2 = \text{Ni} + \text{H}_2\text{O}$, for 2-4 hours at a temperature of 300-400 °C. The process can also be anhydrous, in which it is carried out with a working mixture (methane and water vapor) at 700-800 °C.

If nickel is not in the form of oxides, but in the form of their compounds with aluminum oxide (spinel), then a high temperature (800-900 °C) is required for its recovery. In this case, the recovery process is slow. The formation of H_2 when the nickel-aluminum catalyst is heated above 600 °C.

The GIAP-8 catalyst is designed for the conversion of natural gas. The composition of the catalyst is nickel and aluminum oxide. The mass fraction of NiO in the catalyst is 6-10%.

And in the catalyst used, the nickel oxide content is 5.5-6.5% of the total weight of the catalyst. Given that nickel is purchased abroad at a high price, it is undesirable to throw it away. As a result of scientific research, the extraction of $\text{Ni}(\text{NO}_3)_2$ with nitric acid from the GIAP-8 catalyst used was carried out. As a result, the concentration of acid, its norm, and the ratio of liquid phases are determined.

The resulting $\text{Ni}(\text{NO}_3)_2$ solution was impregnated with a catalyst carrier (alumina-calcium-manganese alloy) CHKR-06 (improved). To extract GIAP-8 from spent catalysts, the method of dissolution in nitric acid was used in the central laboratory of JSC "Maxam-Chirchik". Catalyst for the development of GIAP-8 20%, 30%, 40%, 50% it dissolves in concentrated nitric acid [4]. The results shown in Table-1 show that washing with boiling water after separation in a 40% solution of nitric acid for 4 h. the conversion of 90-95% nickel into solution is achieved (Table 1).

The GIAP-8 catalyst, which was originally used for the process, was crushed and placed in a glass with distilled water in a ratio of 1 to 4 and evaporated by boiling water. History 20%, 30%, 40%, 50% in concentrated nitric acid, insist, stirring for 4 hours. Gives color, slowly passing into the solution for 4 hours. Then the mixture was divided into liquid and solid phases in a vacuum filter. $\text{Ni}(\text{NO}_3)_2$, forming the composition of the liquid phase, was analyzed by titration.

The mass of the spent catalyst GIAP-8, g	Concentration of HNO_3 , %	Amount of acid, amount of water, ml	Amount of water, ml	Amount of NiO remaining in the solid phase, %	Amount remaining in acid, g/dm ³	Degree of extraction Ni, %
100 standard						
10	20	106,75	40	0,967	16,42	82,84
10	30	71,16	40	0,6716	26,53	91,6
10	40	53,375	40	0,42	43,824	95,12
110 standard						
25	20	266,87	-	1,6	14,15	78,27
25	30	178	-	1,56	21,45	79,537
25	40	101,34	-	1,02	40,172	88,75
25	50	106,75	-	0,93	43,824	90,05
25	20	293,56	100	2,328	12,1	72,036
25	30	195,8	100	0,827	21,4	90,33
25	40	146,74	100	0,714	35,38	92,43

Table 1. Spent catalyst GIAP-8 for 4 hours 20%, 30%, 40%, 50% dissolution in concentrated nitric acid

The solid phase was also analyzed for the content of NiO in it. The rate of nickel release from the catalyst was calculated by the formula:

where , C is the mass of nickel oxide isolated from the catalyst (NiO), g ;

C_1 is the mass of nickel oxide (NiO) remaining in the solid phase after filtration of the used GIAP-8, $giap$.

To calculate the degree of dissociation, given that the mass fraction of NiO in the GIAP-8 catalyst used is 6.5%, a mass fraction of 6% NiO was isolated in it. The mass concentration of $\text{Ni}(\text{NO}_3)_2$ in solution is (200-550) g/dm^3 . The mass of nickel oxide remaining in the solid phase is (20-30) g/dm^3 . The nickel nitrate solution was evaporated and the $\text{Ni}(\text{NO}_3)_2$ nickel nitrate content was adjusted to 2-490 g/dm^3 . This solution was immersed twice in the carrier, and the catalyst was dried and cooled in the control mode.

The mass fraction of NiO is 10%, and the mass fraction of SO_3 is 0.004%. CHCR-21 catalyst was obtained. The catalyst test is performed in Figure 1.

Supplementary Files

Figure 1. The scheme of testing the catalyst in the laboratory

The results of the catalyst tests obtained in the central laboratory of JSC "Maxam-Chirchik" are shown in Table 2.

The table of test results of the catalyst ChKR-06 (advanced) test facilities were tested and put into operation. Prepared to work with nitrogen. Hydrogen was given and the temperature began to rise to 500 °C.

Tsh 0020368-15:2014 according to the enterprise standard, the test result of the CHCR-06 (improved) catalyst should not exceed 37% of the residual methane content, which does not decompose at a temperature of 500 °C. At a temperature of 700 °C, the content of non-decomposed residual methane should not exceed 8%. Based on the test results of the catalyst CHKR-06 (advanced), the test result indicators showed that at a temperature of 500 °C, the content of non-decomposed residual methane was 35.7%, and at a temperature of 700 °C, the content of non-decomposed residual methane was 6.5%.

Time	Temperature mode, °C			Water mode			Gas mode	Note
	In the reactor	Sulfur purification	In xylene	Diffmanometer pressure, mm	Measurement of the water level in the burette, ml	Water flow, ml	CH4	
1030	500	250	180	61	19	17	30	Exit №1-35,3%
1100	500	250	180	84,5	23,5	20	30	Exit №2-36,1%
1130	500	250	180	100/6,5	22	17	30	Exit №3-35,8%
1200	600	250	190	25	18,5	11,5	30	
1230	675	250	190	48,5	23,5	13,5	30	
1300	700	250	190	70	21,5	13,5	30	
1330	700	250	185	96	26	15	30	
1400	700	250	185	100/21	25,5	13	30	Exit №1-5,6%
1430	700	250	185	44,5	24	14	30	Exit №2-6,8%
1500	700	250	185	70	24,5	14	30	Exit №3-7,2%
1530	700	250	185	93	23	13	30	Exit №4-6,5%

Table 2. Results of indicators of catalyst activity

Conclusion

The I contained in the GIAP-8 catalyst separated with nitric acid was separated. To do this, water is placed in the catalyst in a ratio of 1: 4 and boiled water is evaporated, then nitric acid is added and infused, stirring for 4 hours. The solid phase was then separated from the liquid phase in a vacuum filter, and the solid phase was analyzed for the $\text{Ni}(\text{NO}_3)_2$ g/dm^3 content in the liquid phase. In the case of $\text{Ni}(\text{NO}_3)_2$ the resulting solution was evaporated and

adjusted to $\text{Ni}(\text{NO}_3)_2 = 490 \text{ g/dm}^3$, and a CHCR-06 (improved) catalyst was installed on the carrier. The impregnation of NIO meets the requirements of GOST. The level of operation of the resulting catalyst in laboratory testing equipment was checked. As a result, the indicators meet the requirements of the enterprise standard.

References

1. F. Ch. Gasanova, V. L. Bagiev "Dependences of the activity of nickel-containing catalysts in the reaction of steam reforming of glycerol on their crystallinity." Problems of modern science and education. - 2016. - No. 7 (49). - P. 12-15.
2. VG Demchenko "Preliminary conversion of methane by recirculation gases". Modern science: research, ideas, results, technologies: Sat. scientific Art. - Kyiv: NPVK "Triacon", 2010. - VIP. 2(4). - S. 201-205.
3. I.P. Kirillova. Scientific basis for the preparation of catalysts. Creative heritage and further development of works: Monograph / Ed. A.P. Or at; GOUVPO Ivan. state chemical-technological un-t. - Ivanovo, 2008. -156 p., ill.
4. I.I.Znachenie, I.I.Kalinichenko, A.N.Gabdullin, A.S. Molodykh, V.E. Zaretsky, O.A. Antropova, E.G. Pecherskikh "Nitric acid technology for the processing of magnesium-containing raw materials for the production of metallic magnesium and its compounds." In the world of scientific discoveries. - Krasnoyarsk, - 2010. - No. 4. - P. 34-36.
5. I. I. Kalinichenko, A. N. Gabdullin, A. S. Molodykh, V. V. Vaitner "Complex methods of nitric acid technology for processing substandard magnesium, nickel and alumina-containing raw materials. Science and technology." XXIX Russian school: abstract. report - Miass: MSNT, - 2009. - P.116.
6. A.S. Molodykh, A.N. Gabdullin, E.A. Nikonenko, V.V. Vaitner "Nitric acid methods of processing substandard raw materials. Collection of materials of the All-Russian student Olympiad, scientific-practical conference and exhibition of works of students, graduate students and young scientists. Energy and resource saving. Non-traditional and renewable energy sources. Yekaterinburg, - 2013. - P.301.
7. Ilkhomovich, F. B., Khujakhmatovich, T. S., & Sabirovich, I. F. Development of Production of Natural Gas Primary Reforming Catalyst. International Journal on Integrated Education, 3(9), 264-266.
8. Farmanov, B. I., Tavashov, Sh. Kh., & Dadakhodzhaev, A. T. (2020). The effect of the amount of the Ca-containing component and the modes of heat treatment of the corundum catalyst support has been studied. in Technical and Technological Modernization of Russia: Problems, Priorities, Prospects (pp. 29-31).
9. Farmanov, B. I., & Tavashov, Sh. H. (2021). Development of a technology for obtaining strong carriers and nickel catalysts for the primary reforming of natural gas. Universum: Engineering Sciences, (5-5), 17-20.
10. Farmanov, B. I., Dadakhodzhaev, A. T., Mirzakulov, H. Ch., & Mingbaeva, D. M. (2020). Technology of catalysts for primary reforming of natural gas. In Chemical Technology and Engineering (pp. 81-82).
11. Ilkhomovich, F. B., & Tursunovich, D. A. (2020). Development of a technology for the production of aluminum-nickel calcium catalyst for steam conversion of natural gas. Asian Journal of Multidimensional Research (AJMR), 9(2), 252-260.
12. Farmanov, B. I., Ruzieva, Z. T., & Farmanova, D. I. (2020). Determination of the optimal parameters for the absorption of carbon dioxide by an ammonia solution of ammonium nitrate. Universum: engineering sciences, (2-2 (71)).